1. $(245 \pm 23) - (22 \pm 2) =$

2. $\frac{100.3 \pm 0.4}{22.1 \pm 0.4} - 5.0 \pm 0.3 =$

3. $\{(12.01 \pm 0.09) + (125.3 \pm 0.1)\}/(11 \pm 1) =$

4. $\frac{1002 \pm 1}{102 \pm 1} =$

5. An unknown solution of NaOH is used to neutralize 20.00 ±0.03 mL of 0.100 ± 0.005 M HCl. The initial buret reading was 5.00 ±0.05 mL. The buret reading at the endpoint was 10.02 ± 0.05 mL. What is the molarity (and error) of the NaOH solution?