7/2.2 8.00 g of \rm NaOHare dissolved in water to make 2.00 L of solution. What is the concentration of hydronium ions, \rm [H_3O^+], in this solution? \rm [H_3O^+] =

Bottom of Form

**7/3.2** Carbonated cola is more acidic than coffee or even orange juice because cola contains phosphoric acid. What is the molar concentration of \rm{H_3O^+}in a cola that has a \rm pHof 2.670?**Express your answer with the appropriate units.** \rm [H_3O^+]=

7/4.2 How many moles of \rm HNO_3are present if 4.70×10−2 mol of \rm Ba(OH)_2was needed to neutralize the acid solution? **Express your answer with the appropriate units.**

**8/2.2** 8.00 g of \rm NaOHare dissolved in water to make 2.00 L of solution. What is the concentration of hydronium ions, \rm [H_3O^+], in this solution?

8/3.1 Carbonated cola is more acidic than coffee or even orange juice because cola contains phosphoric acid. What is the molar concentration of \rm{H_3O^+}in a cola that has a \rm pHof 2.670? **Express your answer with the appropriate units.**

8/4.1 A chemist needs to determine the concentration of a solution of nitric acid, \rm HNO_3. She puts 705 mL of the acid in a flask along with a few drops of indicator. She then slowly adds 0.200 \it M \rm Ba(OH)_2to the flask until the solution turns pink, indicating the equivalence point of the titration. She notes that 235 mL of \rm Ba(OH)_2was needed to reach the equivalence point.

How many moles of \rm HNO_3are present if 4.70×10−2 mol of \rm Ba(OH)_2was needed to neutralize the acid solution? **Express your answer with the appropriate units.**

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