7/2.2 8.00  of are dissolved in water to make 2.00  of solution. What is the concentration of hydronium ions, ![\rm [H_3O^+]](), in this solution? ![\rm [H_3O^+]]() =

Bottom of Form

**7/3.2** Carbonated cola is more acidic than coffee or even orange juice because cola contains phosphoric acid. What is the molar concentration of in a cola that has a of 2.670?**Express your answer with the appropriate units.** ![\rm [H_3O^+]]()=

7/4.2 How many moles of are present if 4.70×10−2  of was needed to neutralize the acid solution? **Express your answer with the appropriate units.**

**8/2.2** 8.00  of are dissolved in water to make 2.00  of solution. What is the concentration of hydronium ions, ![\rm [H_3O^+]](), in this solution?

8/3.1 Carbonated cola is more acidic than coffee or even orange juice because cola contains phosphoric acid. What is the molar concentration of in a cola that has a of 2.670? **Express your answer with the appropriate units.**

8/4.1 A chemist needs to determine the concentration of a solution of nitric acid, . She puts 705  of the acid in a flask along with a few drops of indicator. She then slowly adds 0.200  to the flask until the solution turns pink, indicating the equivalence point of the titration. She notes that 235  of was needed to reach the equivalence point.

How many moles of are present if 4.70×10−2  of was needed to neutralize the acid solution? **Express your answer with the appropriate units.**

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