I have these questions:

1. Equal molar amounts of hydrochloric acid and potassium hydroxide are mixed together in water. Is the resulting solution acidic, basic, or neutral?



2. Equal molar amounts of hypochlorous acid and sodium cyanide are mixed together in water. Is the resulting solution acidic, basic, or neutral? Ka hypochlorous acid = 3.5 x 10-8 ; Ka hydrocyanic acid = 4.9 x 10-10

3. Equal molar amounts of benzoic acid and sodium acetate are mixed together in water. Is the resulting solution acidic, basic, or neutral? Ka benzoic acid = 6.3 x 10-5 ; Ka acetic acid = 1.7 x 10-5

4. Suppose you want to make an acetic acid-sodium acetate buffer solution that maintains a pH of 5.00. How many grams of sodium acetate should be dissolved in 1.00 L of 1.150 M acetic acid? Ka acetic acid = 1.7 x 10-5

5. Which of the following, when added to water, will give a solution that has a pH > 7 : NaClO4 LiNO2 , NiCl2, Sr(NO3)2

