What is the standard Gibbs free energy for the transformation of diamond to graphite?

\rm C_{diamond} \rightarrow C_{graphite}

 {\Delta G_{\rm rxn}}^\circ =-2.90  \rm kJ/mol

**Correct**

http://session.masteringchemistry.com/assets/images/submit_grey.gifhttp://session.masteringchemistry.com/assets/images/hints_grey.gif[my answers](javascript:reviewAnswers(2880823,1);)http://session.masteringchemistry.com/assets/images/showanswer_grey.gif[review part](javascript:reviewPart(2880823,1,'680'))

**Part B**

What can be said about the spontaneity of this reaction?

\rm C_{diamond} \rightarrow C_{graphite}

Top of Form

|  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
|  | |  |  | | --- | --- | |  | The reaction (as written) is spontaneous. | |  | The reverse reaction is spontaneous. | |  | the system is in equilibrium at 298 K. | |  |  |

Bottom of Form