What is the standard Gibbs free energy for the transformation of diamond to graphite?



   =-2.90  

**Correct**



**Part B**

What can be said about the spontaneity of this reaction?



Top of Form

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|  | The reaction (as written) is spontaneous. |
|  | The reverse reaction is spontaneous. |
|  | the system is in equilibrium at 298 K. |

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Bottom of Form