Bottom of Form

**Part B**

The combustion of pentane, , occurs via the reaction



with heat of formation values given by the following table:

|  |  |
| --- | --- |
| Substance | \Delta H^\circ_{\rm f}(\rm kJ/mol) |
| \rm C_5H_{12}(g) | -35.1  |
| \rm CO_2(g) | -393.5  |
| \rm H_2O(g) | -241.8  |

Calculate the enthalpy for the combustion of pentane

**Express your answer in kilojoules using four significant figures.**

|  |  |  |
| --- | --- | --- |
|   \Delta H^\circ_{\rm rxn} = |  |   \rm kJ |
| **Try Again; 4 attempts remaining** |



