The moles for ethanol will be multiplied by 2 because 1 mole of ethanol produces 2 moles of carbon dioxide.

Convert to grams

$mass = n \times M$

$mass = (17.15 \times 2) \times 44$

$mass = 1509.4\ g$

$mass = 1.51\ kg$

$\therefore$ There is $1.51\ kg$ of CO_2 released per litre of ethanol

Moles of water vapor produced

Relative molecular mass:

Mass of hydrogen = 1
Mass of oxygen = 16

$(2 \times 1) + (1 \times 16)$

$= 2 + 16$

$= 18\ g$

The moles for ethanol will be multiplied by 3 because 1 mole of ethanol produces 3 moles of water vapor.

Convert to grams

$mass = n \times M$

$mass = (17.15 \times 3) \times 18$

$mass = 926.2\ g$

$mass = 0.926\ kg$

$\therefore$ There is $0.926\ kg$ of H_2O released per litre of ethanol

$\therefore$ The total amount of gas releasd per litre of ethanol $=$
$1.51 + 0.926$
$= 2.436\ kg$

Change in enthalpy for ethanol:

$\Delta H = energy\ required\ to\ break\ bonds - energy\ released\ when\ new\ bonds\ form$
$= H(reactants) - H(products)$