**Question 1**

Determine the ionic strength, μ, for each of the following solutions. Assume complete dissociation of each salt and ignore any hydrolysis reactions.

1. A solution of 0.00556 M FeSO3

b) A solution of 0.00282 M CuCl2

c) A solution of 0.000853 M CaBr2 and 0.000509 M La(NO3)3

**Question 2**

Part 1

What happens to the solubility of an ionic compound as ionic strength of the solution increases up to ~0.5 M)?

1. Solubility increases
2. Solubility decreases

Part 2

Why is there a correlation between the solubility of an ionic compound and the ionic strength of the solution?

1. The increase in ionic strength increases the concentration of aqueous ion pairs.
2. The increase in ionic strength decreases the concentration of aqueous ion pairs.
3. The increase in ionic strength increases the ionic atmosphere around an individual ion, shielding it from other ions, decreasing the tendency for ions to bind together.
4. The increase in ionic strength increases the ionic atmosphere around an individual ion, increasing the contact between ions, increasing the tendency for ions to bind together.