Find the standard reduction potential for the reaction (it’s a range based off results)

VO3-(aq) + 4H+(aq) + e- 🡪 VO2+(aq) + 2H2O(l) Ecell=??

Results for my reactions any heat that is mentioned was for 12-15 minutes

Vanadyl Sulfate (blue solution):

1.) 20 drops of 3% H2O2: initially turned brown and then pale green after heat

2.) 40 drops of Bromine water: initially green then light blue/green after heat

3.) 20 drops of 0.2M NaNO3: initially blue and after heat blue

4.) 20 drops of Fe2(SO4)3: initially blue and after heat green

5.) 20 drops of 0.2M Na2SO4: initially blue then after heat still blue

6.) control tube

Ammonium Vanadate (yellow solution):

1. pea size amount of amount solid Na2SO3: turned cloudy blue then clear blue
2. 20 drops of distilled water: no reaction and no heat
3. 20 drops of 0.2M NaBr: no reaction and no change after heat
4. 20 drops of 0.2M NaNO2: turned pale green/blue and light blue after heat
5. pea size amount of Fe(NH4)2(SO4)2 \*6H2O: turned turquoise immediately
6. 20 drops of H2C2O4: no reaction initially but turned blue after heat
7. control tube

I have to find the range of the above reaction based on the results we obtained, my lab book does not explain how to go about finding this range and I am at a loss at how to do it.