I know 1% = 10,000 mg/L MG is million gallons

 mg/L lbs lbs

 mil. Lbs. MG

**To express mg/L concentration as lbs/million lbs** requires no concentration conversion at all since mg/L is equivalent to parts per million parts (ppm) concentration. Parts per million can be expressed using any units—milligrams, grams, pounds, etc.

mg/L = 6ppm = mg

 1,000,000 mg

 Or = grams

 1,000,000 grams So, a concentration of 2 mg/L may be stated as:

 2 mg/L = 2 ppm = 2 lbs

 Or = lbs 1,000,000 lbs

 1,000,000 lbs

**To express lbs/MG concentration as lbs/million lbs simply** requires converting million gallons (MG) in the denominator to million lbs, using the conversion factor 8.34 lbs/gal. In this calc you multiply only the denominator of the fraction since you are not changing (or increasing) the amount of water represented. You are merely changing how you are reporting that amt of water (lbs instead of gallons)

Percent Strength Using Dry Hypochlorite

Two Parts of a Solution

1. Solute

Chemical to

Be added

,

 = Solution

1. SOLVENT Xxxxxx or ![C:\Users\Bea\AppData\Local\Microsoft\Windows\Temporary Internet Files\Content.IE5\1MLEALD2\MP900400105[1].jpg]() liquid

 Xxxx

 Xxxxx

 Dry

% = Part x 100 ; Percent chlorine strength is calculated as **part chlorine, in lbs**

 Whole divided by the **whole solution,** in lbs

 % Cl2  Chlorine, lbs x 100

Strength Solution, lbs

**I don’t understand part of this wording** - Since lbs chlorine is equal to (lbs hypochlorite) (% Avail. Cl2 /100), the numerator of the equation can be replaced. And the denominator of the equation (lbs solution) includes both water (lbs) and chlorine or chlorine compound (lbs). Thus the expanded equation may be written as☺

 ( Hypochl, lbs) (% Avail. Cl2)

 % Cl2 = 100 x 100

 Strength Water lbs + Chl Cmpd lbs

As indicated in this last equation, the hypochlorite added must be expressed in pounds. If the hypochlorite weight is expressed in ounces or grams, it must first be converted to pounds.