1. Of the molecules below, select those that are paramagnetic.

It may be useful to draw a molecular orbital energy diagram and to fill in the electrons for each molecule.

Top of Form

Li2B2N2CN−O−2O2−2

1. Give the bond order of each diatomic molecule. Your molecular orbital energy diagrams from the previous question may be useful.





















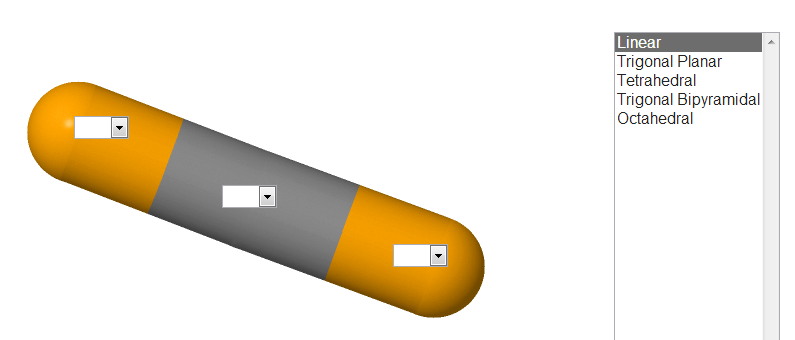


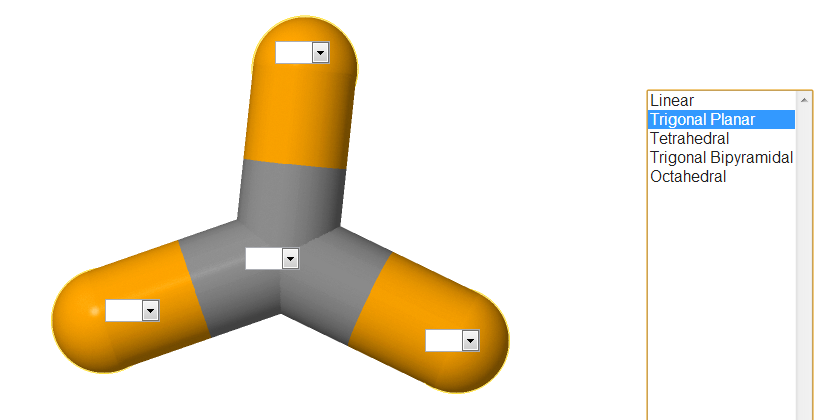


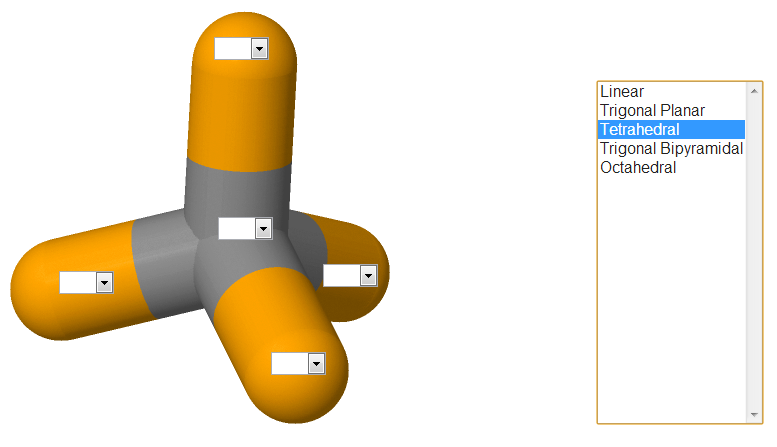
unanswered

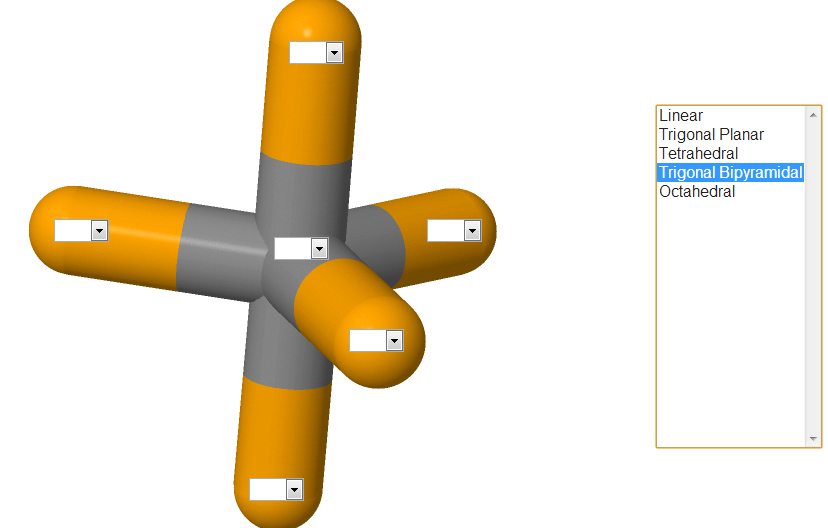
3.) **BF3 is also trigonal planar.**

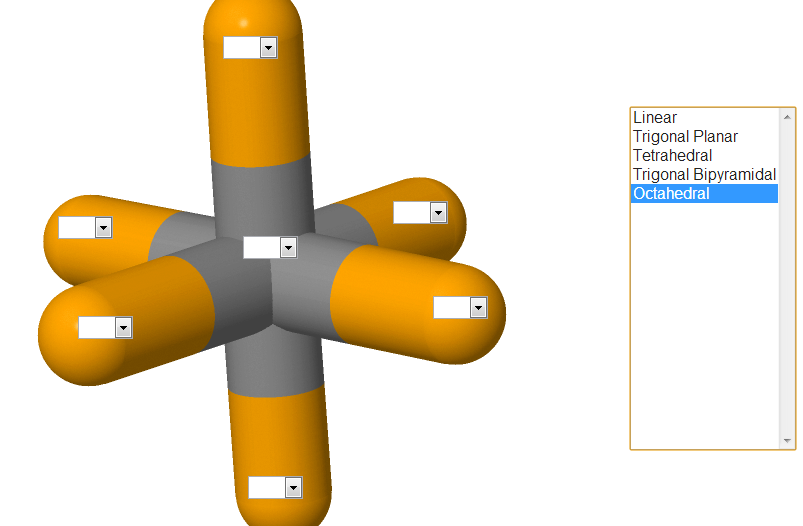
1. Select the appropriate structure for BF3: **B, F por (ep)**
2. Next, place the appropriate atoms into your structure using the internal drop-down menus. If you are using electron pairs in your structure, they can be noted with (ep).Bottom of Form

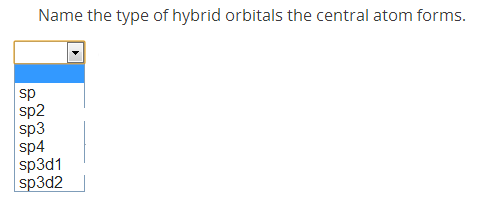


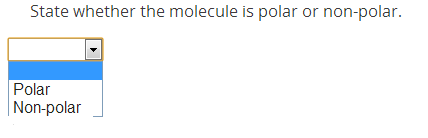




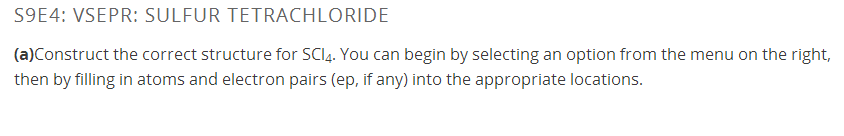


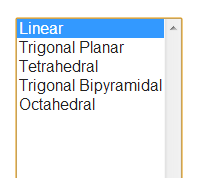


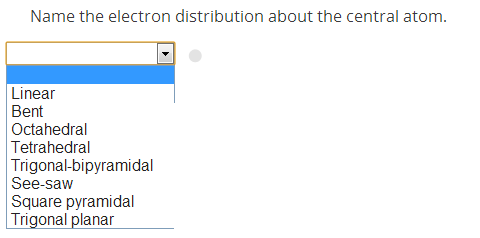
a)

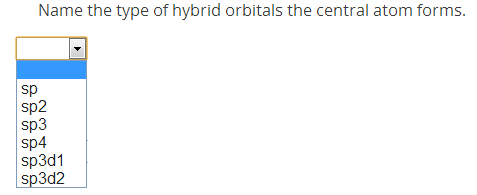


b)

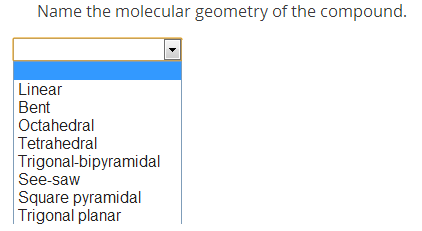
1. 



1. a)



b)

c)

d) State whether the molecule is polar or non-polar.